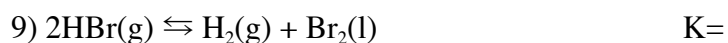
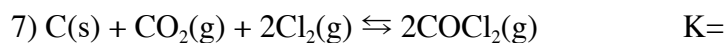
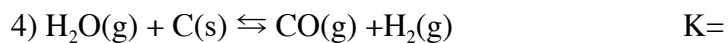
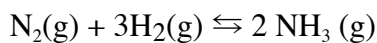


Name _____

Period _____

Equilibrium Constant Practice #2**Part I:** Write the equilibrium constant expression for the following reactions:**Part II:** Calculating equilibrium constants

1) This equation:

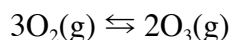


describes a reaction that was carried out at high temperature with the following results:

Trial	N_2	H_2	NH_3
1	0.90	0.600	0.0441
2	0.80	0.500	0.0316
3	0.70	0.400	0.0212

Verify that the equilibrium constant for this is constant by calculating the value for all three trials.

2) This equation:

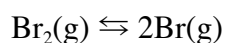


describes a reaction that was carried out at high temperature with the following results:

Trial	O ₂	O ₃
1	0.100	0.0274
2	0.200	0.0775
3	0.300	0.142

Calculate the equilibrium for all three trials.

3) This equation:



describes a reaction that was carried out at 100°C with the following results:

Trial	Br ₂	Br
1	0.250	0.0791
2	0.500	0.112
3	0.750	0.137

Calculate the equilibrium for all three trials.

4) This equation:



describes a reaction that was carried out at high temperature with the following results:

Trial	NOBr	NO	Br ₂
1	0.25	3.95	0.10
2	0.50	5.59	0.20
3	0.75	6.85	0.30

Calculate the equilibrium for all three trials.